

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

Specimen for 2007

GCE A LEVEL

MARK SCHEME

MAXIMUM MARK: 100

SYLLABUS/COMPONENT: 9701/04

**CHEMISTRY
PRACTICAL**



Page 2	Mark Scheme	Syllabus Paper
	GCE A LEVEL – 2007	9701 4

Section A

- 1 (a) (i) 2 x 80g of Br₂ produce 24 dm³ of CO₂ (1)
- Thus 3.2 g of Br₂ will produce $\frac{3.2 \times 24}{2 \times 80} = 0.48 \text{ dm}^3$ (1)
- (ii) Colorimetrically : withdraw samples periodically (1)
measure absorbance (1)
plot absorbance against time (1)
- OR
Iodometrically : oxidising I⁻ to I₂ (1)
by Br₂ (1)
titrating with thiosulphate (1)
- (Allow titration of H⁺ or evolution of CO₂ if some mention of solubility.) [5]
- (b) (i) Reaction has a constant half-life (1)
evidence from graph that t(1/2) is constant (1)
- (ii) Rate = [Br₂] (1)
- (iii) At least two measurements of half-life from first graph (1)
Calculation of mean (say 200 secs) (1)
- [5]
- Total : 10**
- 2 (a) The standard enthalpy change of formation of a compound is the enthalpy change when one mole of a compound is formed (under standard conditions) from its elements in their standard states. (1) [2]
- (b) Suitable cycle clearly labelled showing all three values [2]
- (c) (i) 298 kJ mol⁻¹ (1)
- (ii) In Data Booklet Si-Cl bond energy is 210 kJ mol⁻¹.
SiCl₄ is not a gas under standard conditions (1)
- [2]
- (d) (i) SiCl₃H + H₂ → Si + 3HCl (1)
- (ii) From the Data Booklet, E_{Si-Cl} = 359, E_{H-Cl} = 431, E_{H-Br} = 366 kJ mol⁻¹
Per Si-hal bond, for SiCl₃H, ΔH = 359 - 431 = -72, for SiBr₃H, ΔH = 298 - 366 = -68 (1)
therefore the reaction with SiBr₃H will be less exothermic i.e. overall reaction would be more endothermic (1)
OR ΔH_{reaction}: for SiCl₃H, ΔH_{reaction} = + 96, and for SiBr₃H, ΔH_{reaction} = +108 (1)
therefore overall reaction is more endothermic (1)
- (iii) Manufacture of semiconductors (or equivalent) (1)
- [4]
- Total : 10**

Page 3	Mark Scheme	Syllabus	Paper
	GCE A LEVEL – 2007	9701	4

- 3 (a) $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$ (1)
- $\text{CaO} + \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2$ (1) [2]
- (b) To neutralise acid soils (1)
To improve soil 'quality' by precipitating clays (or equivalent) (1) [2]
- (c) The temperature increases (1)
As the Group is descended, the cation increases in size (1)
Thus ability of the cation to polarise the anion decreases, increasing the stability of the carbonate. (1) [3]
- (d) (i) $\text{CaMg(CO}_3)_2 + 4\text{HCl} \longrightarrow \text{CaCl}_2 + \text{MgCl}_2 + 2\text{CO}_2 + 2\text{H}_2\text{O}$ (1)
- (ii) M_r of dolomite is $40 + 24 + (2 \times 60) = 184$ (1)
- 184 g of dolomite should produce 2×44 g of CO_2
- Hence 1 g of dolomite should give $\frac{88}{184}$ g of $\text{CO}_2 = 0.478$ g
- % purity of the dolomite is $\frac{0.450 \times 100}{0.478} = 94.1\%$ (1) [3]
- Total : 10**
- 4 (a) (i) $[\text{Ar}] 3d^{10} 4s^1$
- (ii) $[\text{Ar}] 3d^{10}$
- (iii) $[\text{Ar}] 3d^9$ [2]
- (b) Any four of the following points:
- colour due to absorption of certain visible frequencies
 - d-orbitals are split into two groups by presence of ligands
 - light absorbed when e^- moves from lower to higher orbital
 - this needs a gap in the higher orbital, so d^{10} in Cu(I) is not coloured
 - if CuCl_2 is blue, then photons absorbed must be red ones
- [4]
- (c) (i) $[\text{Cu(H}_2\text{O)}_6]^{2+} + 2\text{OH}^- \longrightarrow [\text{Cu(H}_2\text{O)}_4(\text{OH})_2] + 2\text{H}_2\text{O}$
 $(\text{NH}_3 + \text{H}_2\text{O} \longrightarrow \text{NH}_4^+ + \text{OH}^-)$
 $[\text{Cu(H}_2\text{O)}_4(\text{OH})_2] + 4\text{NH}_3 \longrightarrow [\text{Cu(NH}_3)_4]^{2+} + 2\text{OH}^- + 4\text{H}_2\text{O}$
 (or $\rightarrow [\text{Cu(NH}_3)_4(\text{H}_2\text{O})_2]^{2+}$) (2)
- (ii) These are *ligand exchange* reactions.
 H_2O is exchanged for OH^- , and H_2O and OH^- are exchanged for NH_3 . (2) [4]
- [Total : 10]**

Page 4	Mark Scheme	Syllabus	Paper
	GCE A LEVEL – 2007	9701	4

- 5 (a) Van der Waals forces increase with the number of electrons present (1)
 Since this allows larger dipoles and hence stronger attractive forces (1) [2]
- (b) (i) Description of Cl^- , Br^- , and I^- with conc sulphuric acid 3 x (1)
 Use of E^\ominus (1)
- (ii) Description of HCl , HBr and HI 3 x (1)
 Use of E^\ominus (1) [8]

Total : 10

6 (a)

element	%	A_r	% / A_r	ratio
C	40.0	12	3.33	1
H	6.65	1	6.65	2
O	53.3	16	3.33	1

[1]

- (b) (i) It contains an asymmetric carbon atom (1)
- (ii) It contains a carboxylic acid group (1)
- (iii) It contains a $CH_3CH(OH)-$ or CH_3CO- group (1) [3]
- (c) Displayed formula of 2-hydroxypropanoic acid [1]
- (d) Displayed formula of the ketone of the above [1]
- (e) Displayed formula of the cyclic di-ester (1)
 Ester (1) [2]
- (f) Displayed formula of 3-hydroxypropanoic acid (1)
 Compound C is $CH=CHCO_2H$ (1) [2]

Total : 10

Page 5	Mark Scheme	Syllabus	Paper
	GCE A LEVEL – 2007	9701	4

7	(a) (i)	1. amine 3. amide	2. carboxylic acid 4. ester	(2)	
	(ii)	1. amine Both can form ions (both polar groups) by gain or loss of a proton from water <u>or</u> form hydrogen bonds with water	2. carboxylic acid	(1)	
	(b) (i)	Allow conc. HCl <u>and</u> heat or conc. NaOH <u>and</u> heat		2 x (1)	
	(ii)	Diagrams of aspartic acid, phenylalanine and methanol		3 x (1)	[5]
	(c)	It could be decomposed/hydrolysed during cooking			[1]
					Total :10

Section B

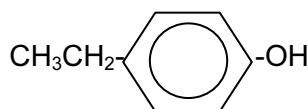
8	(a)	6 points from the following :			
		<ul style="list-style-type: none"> • 2 strands of DNA separate • mRNA reads the 'code'/base sequence on the DNA • mRNA moves out of the nucleus • mRNA binds to the ribosome • tRNA binds to amino acids • amino acids are transferred to ribosome and joined to growing chain • until Stop codon is reached 			[6]
	(b)	Each amino acid needs 3 bases to code for it		(1)	
		3 x 129 = 387, which leaves 3 bases to code for Start and 3 for Stop		(1)	[2]
	(c) (i)	A variety of answers possible e.g. <ul style="list-style-type: none"> • sickle cell disease • thalassaemia • cystic fibrosis • haemophilia etc. 		(1)	
	(ii)	A suitable symptom e.g. <ul style="list-style-type: none"> • Deformed red blood cells • Restricts production of haemoglobin • Mucous lining of lungs thickens • Poor clotting of blood/bleeding under the skin 		(1)	[2]
					Total : 10

Page 6	Mark Scheme	Syllabus Paper
	GCE A LEVEL – 2007	9701 4

9 (a)	%	%/A _r	Ratio				
	C 78.7	6.56	8)	Empirical formula		
	H 8.2	8.2	10) (1)	C ₈ H ₁₀ O	(1)	
	O 13.1	0.82	1)			
	<i>M_r</i> = 122, hence this molecular formula						
	Molecular formula is C ₈ H ₁₀ O					(1)	[3]


- (b) 1.2δ - CH₃ (1)
- 2.5δ - CH₂ (1)
- 5.5δ - OH (1)
- 6.8δ aryl hydrogens x 4 (1)

Hence structure is



(or ethyl phenol isomers)

- (c) Peak at 5.5δ would disappear (1)
- Due to rapid exchange with D⁺ which does not absorb here (1) [2]

- (d) CH₃OCH₂- or isomers (1)

Two sensible suggestions (2) [3]

Total : 12

- 10 (a) Can be used as a fuel (for generating electricity) (1)
- Can be hydrolysed (using acid or enzymes) and the sugars fermented (1) [2]

(b) Carbon dioxide [1]

(c) (C₆H₁₀O₅)_n + nH₂O → nC₆H₁₂O₆ [1]

- (d) Ethanol has an -OH group and so can be washed away
Gasoline is a hydrocarbon and is not soluble in water
Gasoline requires detergent which can add to the pollution
Ethanol is biodegradable (any 3) [3]

Total : 7