

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the October/November 2011 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/22

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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- 1 (a) (i) mass of C = $\frac{12 \times 0.352}{44} = 0.096\text{g}$ (1)
- $n(\text{C}) = \frac{0.096}{12} = 0.008$ (1)
- (ii) mass of H = $\frac{2 \times 0.144}{18} = 0.016\text{g}$ (1)
- $n(\text{H}) = \frac{0.016}{1} = 0.016$ (1)
- (iii) mass of oxygen = $0.240 - (0.096 + 0.016) = 0.128\text{g}$ (1)
- $n(\text{O}) = \frac{0.128}{16} = 0.008$ (1)
- allow ecf at any stage [6]
- (b) C : H : O = 0.008 : 0.016 : 0.008 = 1:2:1
- allow C : H : O = $\frac{0.096}{12} : \frac{0.016}{1} : \frac{0.128}{16} = 1:2:1$
- gives CH₂O (1) [1]
- (c) (i) $M_r = \frac{mRT}{pV} = \frac{0.148 \times 8.31 \times 333}{1.01 \times 10^5 \times 67.7 \times 10^{-6}}$ (1)
- = 59.89
- allow 59.9 or 60 (1)
- (ii) C₂H₄O₂ (1) [3]
- (d) CH₃CO₂H (1)
- HCO₂CH₃ (1) [2]
- (e) the only products of the reaction are the two oxides H₂O and CO₂ and copper (1) [1]

[Total: 13]

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- 2 (a) $S(g) \rightarrow S^+(g) + e^-$
 correct equation (1)
 correct state symbols (1) [2]

- (b) from Na to Ar,
 electrons are added to the same shell/have same shielding (1)
 electrons are subject to increasing nuclear charge/proton number (1)
 electrons are closer to the nucleus **or** atom gets smaller (1) [3]

- (c) (i) **Mg and Al**
 in Mg outermost electron is in 3s **and**
 in Al outermost electron is in 3p (1)

3p electron is at higher energy **or**
 is further away from the nucleus **or**
 is more shielded from the nucleus (1)

- (ii) **S and P**
 for S one 3p orbital has paired electrons **and**
 for P 3p sub-shell is singly filled (1)

paired electrons repel (1) [4]

- (d) (i) and (ii)

element	Na	Mg	Al	Si	P	S
conductivity	high	high	—	moderate	low	low
melting point	low	high	—	high	low	low

(1) (1) (1) (1) (1)

one mark for each correct column [5]

- (e) germanium/Ge (1) [1]

[Total: 15]

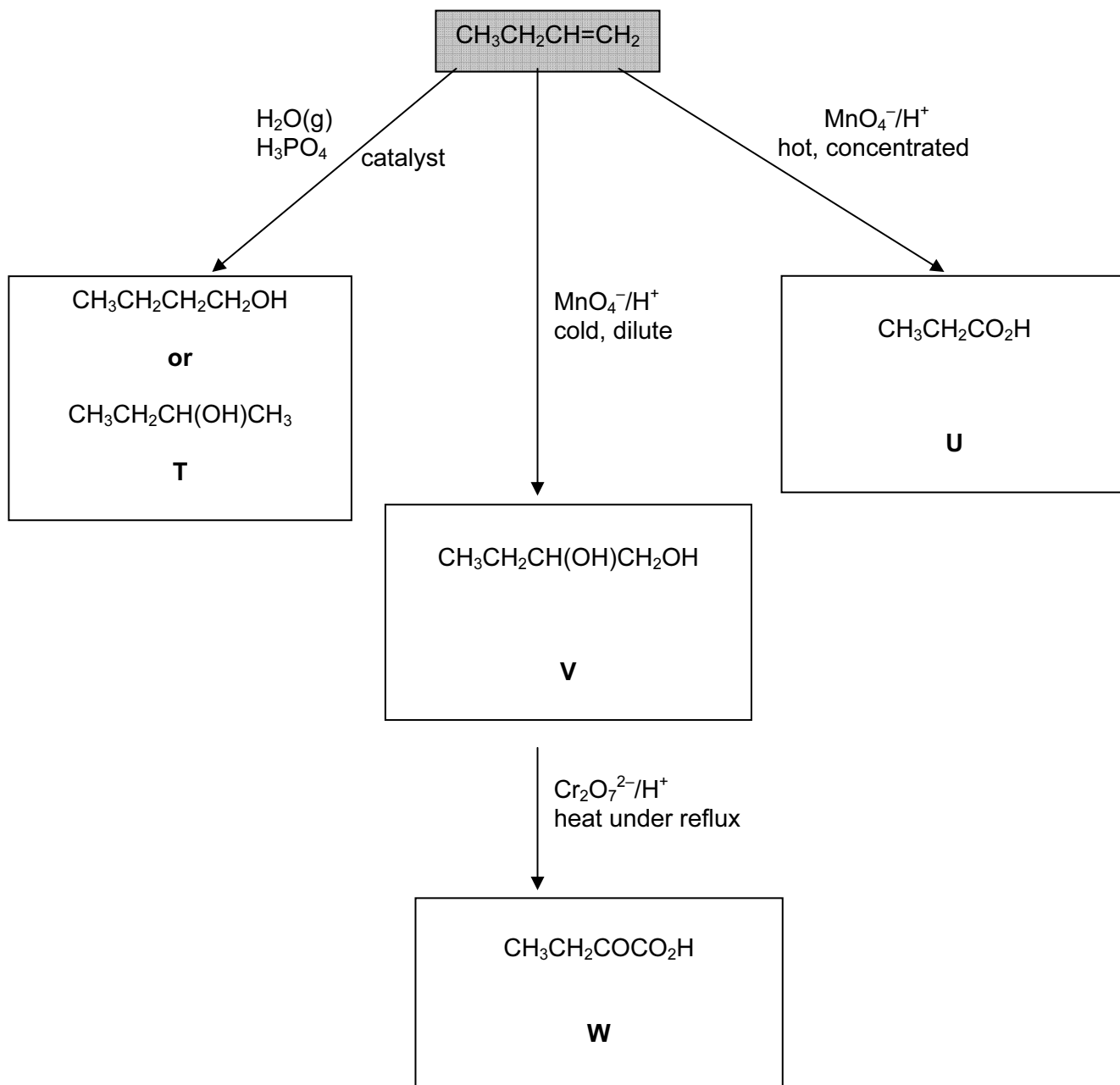
Page 4	Mark Scheme: Teachers' version	Syllabus	Paper 1
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- 3 (a) the overall enthalpy change/energy change/ ΔH for a reaction (1)
- is independent of the route taken **or**
 is independent of the number of steps involved
 provided the initial and final conditions are the same (1) [2]
- (b) (i) $K_2CO_3 + 2HCl \rightarrow 2KCl + H_2O + CO_2$ (1)
- (ii) heat produced = $m \times c \times \delta T = 30.0 \times 4.18 \times 5.2$
 $= 652.08 \text{ J per } 0.0200 \text{ mol of } K_2CO_3$ (1)
- (iii) $0.020 \text{ mol } K_2CO_3 \equiv 652.08 \text{ J}$
 $1 \text{ mol } K_2CO_3 \equiv \frac{652.08 \times 1}{0.0200} = 32604 \text{ J}$
 enthalpy change = $-32.60 \text{ kJmol}^{-1}$ (1)
- (iv) to prevent the formation of $KHCO_3$ **or**
 to ensure complete neutralisation (1) [4]
- (c) (i) $KHCO_3 + HCl \rightarrow KCl + H_2O + CO_2$ (1)
- (ii) heat absorbed = $m \times c \times \delta T = 30.0 \times 4.18 \times 3.7$
 $= 463.98 \text{ J per } 0.0200 \text{ mol of } KHCO_3$ (1)
- (iii) $0.020 \text{ mol } KHCO_3 \equiv 463.98 \text{ J}$
 $1 \text{ mol } KHCO_3 \equiv \frac{463.98 \times 1}{0.0200} = 23199 \text{ J}$
 enthalpy change = $+23.20 \text{ kJmol}^{-1}$ (1) [3]
- (d) $\Delta H = 2 \times (+23.20) - (-32.60) = +79.00 \text{ kJ mol}^{-1}$ (2) [2]

[Total: 11]

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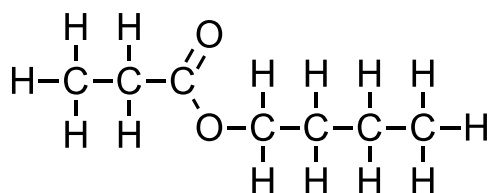
4 (a)



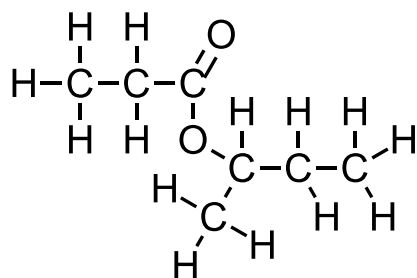
correct **T**
 correct **U**
 correct **V**
 correct > CO group in **W**
 correct -CO₂H group in **W**

(1)
 (1)
 (1)
 (1)
 (1) [5]

(b) T + U



or



correct structures
correctly displayed ester group

(1)
(1) [2]

[Total: 7]

- 5 (a) (i) 1 primary alcohol **not** hydroxyl (1)
alcohol **not** hydroxyl (1)
2 aldehyde **not** carbonyl (1)

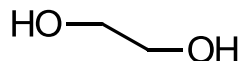
(ii)

test 1			
reagent	Na	$\text{PCl}_3/\text{PCl}_5/\text{PBr}_3$	$\text{RCO}_2\text{H}/\text{H}^+$
observation	gas/ H_2 /effervescence/ fizzing	$\text{HC}\ell/\text{HBr}$ steamy fumes	fruity smell
test 2			
reagent	Tollens' reagent	Fehling's reagent	2,4-dinitro- phenylhydrazine
observation	Ag mirror/silver/ black ppt	brick-red ppt red ppt	orange/red/yellow ppt/solid

only award the observation mark if reagent is correct

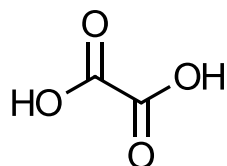
(4) [7]

(b) (i)



(1)

(ii)



(1) [2]

5 (c)

route	starting compound	first reagent	intermediate X	second reagent	intermediate Y	third reagent	final compound
A/1	HOCH ₂ CHO	PCl ₃ PCl ₅ SOCl ₂ etc.	ClCH ₂ CHO	K ₂ Cr ₂ O ₇ /H ⁺ KMnO ₄ /H ⁺ KMnO ₄ /OH ⁻ Tollens' or Fehling's reagents	ClCH ₂ CO ₂ H	NH ₃	H ₂ NCH ₂ CO ₂ H
A/2	HOCH ₂ CHO	HBr P/Br ₂ etc.	BrCH ₂ CHO	K ₂ Cr ₂ O ₇ /H ⁺ KMnO ₄ /H ⁺ KMnO ₄ /OH ⁻ Tollens' or Fehling's reagents	BrCH ₂ CO ₂ H	NH ₃	H ₂ NCH ₂ CO ₂ H
B/1	HOCH ₂ CHO	PCl ₃ PCl ₅ SOCl ₂ etc.	ClCH ₂ CHO	NH ₃	H ₂ NCH ₂ CHO	K ₂ Cr ₂ O ₇ /H ⁺ KMnO ₄ /H ⁺ KMnO ₄ /OH ⁻ Tollens' or Fehling's reagents	H ₂ NCH ₂ CO ₂ H
B/2	HOCH ₂ CHO	HBr P/Br ₂ etc.	BrCH ₂ CHO	NH ₃	H ₂ NCH ₂ CHO	K ₂ Cr ₂ O ₇ /H ⁺ KMnO ₄ /H ⁺ KMnO ₄ /OH ⁻ Tollens' or Fehling's reagents	H ₂ NCH ₂ CO ₂ H
C	HOCH ₂ CHO	Tollens' or Fehling's reagents	HOCH ₂ CO ₂ H	KBr/conc. H ₂ SO ₄	BrCH ₂ CO ₂ H	NH ₃	H ₂ NCH ₂ CO ₂ H
mark		(1)	(1)	(1)	(1)	(1)	

[5]

[Total: 14]