

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the October/November 2010 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/34

Paper 3 (Advanced Practical Skills), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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Question	Sections	Indicative material	Mark
1 (a)	PDO layout	I Volume given for Rough titre and accurate titre details tabulated.	1
	MMO Collection	II In the correct spaces, records initial and final burette readings for Rough titre and ; Initial and final burette readings and , volume of FB 2 added recorded for each accurate titre <i>Headings should match readings.</i> <i>Do not award this mark if:</i> <i>50(.00) is used as an initial burette reading;</i> <i>More than one final burette reading is 50.(00);</i> <i>Any burette reading is greater than 50.(00)</i>	1
	MMO Decisions	III Has two uncorrected, accurate titres within 0.1 cm ³ <i>Do not award this mark if having performed two titres within 0.1 cm³ a further titration is performed which is more than 0.10 cm³ from the closer of the initial two titres, unless a fourth titration, within 0.1 cm³ of the third titration or of either of the pair has also been carried out.</i>	1
	PDO Recording	IV All accurate burette readings (initial and final) recorded to nearest 0.05 cm ³ . <i>Assessed on burette readings only.</i>	1
	MMO Quality	V, VI and VII Round any burette readings to the nearest 0.05 cm ³ . Check and correct subtractions in the titre table. Select the “best” titre using the hierarchy: two identical; titres within 0.05 cm ³ , titres within 0.10 cm ³ etc. Award <u>V, VI and VII</u> for a difference to Supervisor within 0.15 cm ³ Award <u>V and VI only</u> for a difference of 0.15+ cm ³ – 0.25 cm ³ Award <u>V only</u> for a difference of 0.25+ cm ³ – 0.40 cm ³ <i>If the selected “best” titres are > 0.40 cm³ apart, cancel one of the Q marks awarded.</i>	3

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(b)	ACE Interpretation	<p>Calculates the mean, correct to 2 decimal places (third decimal place rounded to the nearest 0.05 cm³) from any accurate titres within 0.20 cm³. <i>A mean of exactly .x25 or .x75 is allowed but the candidate may round up or down to the nearest 0.05 cm³.</i> <i>If ALL burette readings are given to 1 decimal place then the mean can be given to 1 decimal place if numerically correct without rounding.</i> <i>Mean of 24.3 and 24.4 = 24.35 (✓)</i> <i>Mean of 24.3 and 24.4 = 24.4 (✗)</i> Titres to be used in calculating the mean must be clearly shown – in an expression or ticked in the titration table.</p>	1	[1]
(c)	<p>ACE Interpretation</p> <p>PDO Display</p>	<p>No additional factor/expression is allowed in any step <i>If an answer, with no working, is given in any section allow if correct.</i></p> <p>I Uses $\frac{2.00}{158.0}$ in step (i) and answer (i) $\times \frac{\text{cand titre}}{1000}$ in step (ii)</p> <p>II Uses answer (ii) $\times 5$ in step (iii) and answer (iii) $\times \frac{1000}{25}$ in step (iv)</p> <p>III Uses answer (iv) $\times 151.9$ in step (v), and answer (v) $\times \frac{100}{21.50}$ in step (vi)</p> <p>IV Appropriate working shown in a minimum of four sections.</p> <p>V 3 to 5 significant figures in final answers to all sections attempted – minimum of four final answers required</p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p>	[5]
			[Total: 13]	

2	(a)	PDO Layout	I	Records at least four different balance readings and at least one mass of solid/gas <i>Accept 0.0(0X) g as the mass of the empty tube or a statement that the tube is tared.</i>	1	
		PDO Recording	II	Gives all appropriate headings and units when recording results. <i>Do not accept mass of empty tube as 0.0(00)g here unless tube is described as tared. (minimum of three pieces of information)</i>	1	
			III	All recorded balance readings consistent to at least 1 decimal place. <i>(minimum of three balance readings)</i>	1	
		MMO Decisions	IV	Evidence of reheating to “constant” mass. For balances reading to 1 d.p. two masses must be identical For 2 or 3 d.p. balances, two masses must be within 0.05 g	1	
		MMO Quality	V and VI	checks and corrects if necessary all subtractions in the results table. Calculate $\frac{\text{mass heated}}{\text{mass of residue}}$ to 3 significant figures. Compare to supervisor standard or standard value of 1.40. Award <u>V and VI</u> for a difference up to 0.10 Award <u>V only</u> for a difference of 0.10+ to 0.20 <i>Where a candidate repeats the experiment use cumulative masses of FB 3 and residue. Where masses of FB 3 and residue cannot be checked, accept candidate values to calculate the ratio.</i>	2	
	(b)	ACE Interpretation	(i)	Calculates 2.71, (2.710, 2.7097) and	1	
		ACE Conclusions	(ii)	Has: cand value in (i) x mass loss from table in (a) If no mass loss is recorded in the table, check the value used.	1	
			(iii)	Ticks the appropriate box for the experiment and makes some comparison between mass of NaHCO ₃ and the mass of FB 3 used If mass of NaHCO ₃ calculated in (ii) ≥ mass of FB 3 , ignore any ticked box but award the mark for any statement that the mass is not possible.	1	[6]
						[2]

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(c)	ACE Improvements	<p>(i) No mass change with Na_2CO_3 (on heating).</p> <p>(ii) Evidence for no gas produced, e.g.: limewater unaffected, no gas collected in a gas syringe</p> <p><i>If there is reference to measuring mass and to measuring volume but the absence of change is not mentioned, award one of the two marks available.</i></p>	1 1	[2]
(d)	ACE Interpretation	<p>Max errors of 0.05, 0.005 and 0.0005 respectively for balances A, B and C.</p> <p>Calculates:</p> <p>1.11% error for balance A 0.25% error for balance B 0.20% error for balance C</p> <p>Allow ecf on % errors only if:</p> <p>(i) <i>Max errors given are 0.1, 0.01 and 0.001 respectively for balances A, B and C and % errors are 2.22%, 0.50% and 0.40%</i></p> <p>(ii) <i>All max errors are incorrect by a factor 10 e.g. 0.5, 0.05 and 0.005. % errors are 11.1%, 2.5% and 2.0%</i></p>	1 1	[2]
[Total: 12]				

FB 4 is MnSO ₄ (aq); FB 5 is MgSO ₄ (aq); FB 6 is Al ₂ (SO ₄) ₃ (aq); FB 7 is (NH ₄) ₂ SO ₄ (aq)			
3	(a)	MMO Collection	<p>Give one mark for each of the following:</p> <p>I for FB 4 – tests (i) and (iv) 1</p> <p>II for FB 5 – tests (i) and (iv) 1</p> <p>III for FB 6 – tests (i) and (iv) 1</p> <p>IV for FB 7 – tests (i), (iii) and (iv) 1</p> <p>V Give one mark for any change/darkening of the initial precipitate in test (ii) for FB 4 to a qualified brown. 1 <i>The darkening may be described in test (i) or in test (iv)</i></p> <p>VI Describes the test on gas for ammonia in test (iii) for any solution that has no precipitate in either part test of (i) and is warmed. 1 <i>The test for ammonia is expected with FB 7</i> <i>Do not award (VI) if the test is carried out with a solution in which a precipitate had formed at any stage</i> or <i>If a solution in which no precipitate is formed is not warmed with sodium hydroxide</i></p>
			[6]

Results required with NaOH(aq) and NH₃(aq) for the award of marks I to IV in 3(a)

test		observations			
		FB 4	FB 5	FB 6	FB 7
(i)	addition of NaOH	off-white, pale brown, buff or beige precipitate <i>Do not accept cream or equivalent colour precipitates</i>	white precipitate	white precipitate	No precipitate or no change <i>Do not accept clear on its own as an observation; clear solution is acceptable</i>
	further addition of NaOH	precipitate insoluble	precipitate insoluble	precipitate soluble	no precipitate or no change <i>(may be left blank)</i>
(iii)	warming solution with NaOH				any reference to a gas being evolved or reference to red litmus turning blue
(iv)	addition of NH ₃	as NaOH	as NaOH	as NaOH	as NaOH
	further addition of NH ₃	as NaOH	as NaOH	precipitate insoluble	as NaOH

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(b)	ACE Conclusions	<p>Do not accept any ion other than Mn^{2+}, Mg^{2+}, Al^{3+} or NH_4^+ in any section.</p> <p>Marks I to III</p> <p>Ions must be correct, including charge, if a symbol has been given. – <u>no ecf in this section.</u></p>		
		<p>Award <u>I only</u> if one ion only is identified from correct observations.</p>	1	
		<p>Award <u>I and II</u> if two ions only are identified from correct observations.</p>	1	
		<p>Award <u>I, II and III</u> if all four cations are identified from correct observations. <i>The 4th cation may be identified by elimination from incomplete supporting evidence.</i></p> <p>A deduction of Mn^{2+} is allowed from a cream ppt with NaOH(aq) and NH_3(aq)</p>	1	
		<p>IV Award this mark if the supporting evidence fits the ion identified and the practical performed for at least three of the four ions</p> <p><u>Allow ecf on ion order for mark IV.</u> $(Mg^{2+}$ and Al^{3+} are most likely to be interchanged depending on “solubility in excess” observations.</p>		[4]

Minimum evidence required in observations for the ion identity marks I, II and III.

In some cases, identification may be allowed from incomplete observations. There must, however, be no observations that are contrary to those expected with any “correctly” identified ion.

The same criteria will be applied to “candidate’s supporting evidence in awarding mark IV.

Candidates are not permitted to introduce (from the Qualitative Analysis Notes) supporting evidence that is not given in the observations.

Mn^{2+}	off-white precipitate with each reagent, or off-white precipitate turning brown with either of the reagents identification of the ion is allowed from an incorrect observation of a cream or yellow-white precipitate – one ion is known to be Mn^{2+}
Mg^{2+}	white precipitate, insoluble in (excess) NaOH
Al^{3+}	white precipitate, soluble in (excess) NaOH
NH_4^+	no precipitate/no change with either reagent or ammonia, alkaline gas or gas turning red litmus blue evolved

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(c)	MMO Collection	Records no precipitate/no reaction with each of the reagents.	1	[1]
(d)	ACE Conclusions	States that Pb^{2+} /lead(II) would give similar results. <i>Award this mark providing there are no contrary observations for the solution identified as containing Al^{3+}</i>	1	[1]
(e)	MMO Collection	Records a white ppt in (i) Records a yellow precipitate or precipitate turning yellow in (ii).	1 1	[2]
(f)	ACE Conclusions	Award one mark for any attempt to describe replacement of Cl by I in the ppt.	1	[1]
			[Total: 15]	