

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

MARK SCHEME for the October/November 2007 question paper

9701 CHEMISTRY

9701/31

Paper 31 (Practical 1), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the October/November 2007 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



Page 2	Mark Scheme	Syllabus	Paper 1
	GCE A/AS LEVEL – October/November 2007	9701	

Generic Mark Scheme

Skill		Breakdown of marks	
Manipulation, measurement and observation	16 marks	Successful <u>collection</u> of data and observations	8 marks
		<u>Decisions</u> relating to measurements or observations	8 marks
Presentation of data and observations	12 marks	<u>Recording</u> data and observations	5 marks
		<u>Display</u> of calculation and reasoning	3 marks
		Data <u>layout</u>	4 marks
Analysis, conclusions and evaluation	12 marks	<u>Interpretation</u> of data or observations and identifying sources of error	6 marks
		Drawing <u>conclusions</u>	5 marks
		Suggesting <u>improvements</u>	1 mark

Page 4	Mark Scheme	Syllabus	Paper 1
	GCE A/AS LEVEL – October/November 2007	9701	

Question	Sections	Indicative material	Mark	
	<p>Accuracy Calculate (vol of FA 1 x time) for expt 1 and the two additional expts with greatest volume of FA 1. (<i>Round all times to the nearest second</i>) Record the Vt values against the appropriate expt on the candidate's script.</p>			
1 (b) contd.	MMO Decisions	<p>(viii) and (ix) Award both of these marks if two of the Vt values are within 10% of the larger of the closest pair. <i>[Award point (ix) but not point (viii) for a difference of 10+% to 20%]</i></p>	2	[11]
		<p>(x) and (xi) Award both of these marks if candidate's time for expt 1 is within 10% of that obtained by the Supervisor. <i>[Award point (xi) but not point (x) for a difference of 10+% to 20%]</i></p>	2	
<p>Where experiment 1 has been repeated, assess accuracy using the time on page 3. Use the value on page 4 when checking the graph.</p>				
1 (c)	PDO Layout	<p>Ignore labels – check which numerical values have been plotted Ignore omission of negative signs; direction of numbers on axes etc.</p>		
		<p>Plots a rate ($1/t$ or $(\log 1/t)$) on y-axis and a concentration (volume of FA 1 or $(\log \text{volume of FA 1})$) on x-axis <i>If labels correct but numbers on scale indicate a different quantity do not award this mark</i></p>	1	
		<p>Easy to use scales chosen with plotted points covering more than $\frac{1}{2}$ of each available axis</p> <p>A point must be plotted for each experiment performed – <i>take care where expts 1&2 have been omitted from the main results table</i> All points plotted to within $\frac{1}{2}$ small square and in the correct half of a small square</p>	1	

Page 6	Mark Scheme	Syllabus	Paper 1
	GCE A/AS LEVEL – October/November 2007	9701	

Question	Sections	Indicative material	Mark	
1 (e)	ACE Interpretation	<p>Explains that uncertainty is less when change is rapid (or converse)</p> <p>Estimated errors expt 1 number of seconds \leq 3 s. expt 2 number of seconds \geq 3 s, up to a maximum of 10 s. error for expt 2 > error for expt 1</p> <p>Candidate's uncertainties correctly expressed as % of reaction time. Error may be carried forward.</p>	1 1 1	[3]
1 (f)	ACE Improvements	<p>Has: constant volumes of FA 1, variable volume of FA 2, water to keep total volume constant at 55 cm³ Record the total volume for each experiment to the left of the table.</p>	1	[1]
1 (g)	PDO Display	<p>Uses experimental data to make appropriate comment, from experimental results, as to how rate varies with acid concentration. <i>[Do not give this mark where mixtures selected in (f) are not appropriate]</i></p> <p><i>Little change in reaction time is expected. The rate increases slightly as the concentration of acid is doubled etc.</i></p> <p><i>Where an acceptable qualitative statement has been given ignore any incorrect attempt at a quantitative/mathematical expression.</i></p>	1	[1]
			Qn 1 Total 25	

Page 7	Mark Scheme	Syllabus	Paper 1
	GCE A/AS LEVEL – October/November 2007	9701	

Question	Sections	Indicative material	Mark	
<p>FA 3 is aqueous ammonium sulphite (<i>ammonium chloride + sodium sulphite</i>), FA 4 is aqueous sodium nitrite, FA 5 is aqueous barium nitrate, FA 6 is aqueous manganese(II) sulphate,</p>				
2 (a)	MMO Decisions	Chooses named dilute acid as single reagent for identifying nitrite and chooses BaCl ₂ / Ba(NO ₃) ₂ (Ba ²⁺ (aq) or aqueous solution containing Ba ²⁺ acceptable) together with HCl/HNO ₃ (not H ₂ SO ₄) as reagents for identifying sulphate/sulphite	1	[1]
2 (b)	MMO Collection	<p>Award the C3 marks only from observations in the table. No retrospective marks.</p> <p>Give one mark for a brown gas evolved from FA 4 with any acid</p> <p>Give one mark for one of the following for FA 6:</p> <ol style="list-style-type: none"> a white ppt with Ba²⁺ insoluble in hydrochloric or nitric acid, a white ppt with Ba²⁺ insoluble in unnamed acid, precipitate whose colour has not been described insoluble in named acid other than H₂SO₄ <p>Give one mark for one of the following for FA 3:</p> <ol style="list-style-type: none"> a white ppt with Ba²⁺ soluble in hydrochloric or nitric acid, a white ppt with Ba²⁺ soluble in unnamed acid, precipitate whose colour has not been described soluble in named acid other than H₂SO₄ <p><i>If sulphuric acid is <u>stated in the table in (b)</u> award one observation mark only, if the barium salt is added <u>before</u> the acid and a white ppt is obtained with FA 6 and with FA3.</i></p> <p>OR</p> <p>Give one mark for adding HCl/HNO₃ and <u>detecting SO₂</u> (dichromate turning green) from FA 3</p> <p>Give one mark for adding Ba²⁺(aq) and observing a white ppt with FA 6 and no ppt with FA 3.</p> <p><i>Do not give the first of these marks if dichromate is added to the acidified mixture (additional reagent) but allow conclusions from that test)</i></p>	1 1 1	(1) (1)

Page 8	Mark Scheme	Syllabus	Paper 1
	GCE A/AS LEVEL – October/November 2007	9701	

Question	Sections	Indicative material	Mark	
2 (b) contd	ACE Conclusions	Give one mark each if the ions are correctly identified (No evidence is required in this section for the identification of the ions) Mark conclusions consequentially to observations FA 4 – nitrite FA 6 – sulphate FA 3 – sulphite		1 1 1
	ACE Interpretation	Gives appropriate evidence for identification of two of the ions. This mark can be awarded for <u>correct</u> anions where no observations are tabulated.		1 [7]
2 (c)	MMO Collection	FA 6 – observes off-white ppt insoluble in excess with both NaOH and NH ₃ (aq)	1	
		FA 5 – observes white ppt in (iii) and white ppt soluble in HCl in (iv) <i>Ignore any ppt with NaOH/NH₃.</i>	1	
		FA 3 – positive test for alkaline gas described with FA 3 only .	1	[3]
2 (d)	ACE Conclusions	Give two marks if all ions are correctly identified. FA 6 – Mn ²⁺ (from off-white ppt with NaOH and NH ₃ (aq)) or off white ppt with NaOH or NH ₃ (aq) soluble in excess) FA 5 – Ba ²⁺ (if observation mark given in (c) or white ppt with H ₂ SO ₄ and no ppt with NH ₃ (aq)) FA 3 – NH ₄ ⁺ (from ammonia/alkaline gas on warming with NaOH) <i>Give one of these two marks for 2 correct ions Mark conclusions consequentially to observations, e.g. Fe³⁺ from brown ppt for FA 6. Do not allow Ca²⁺ or Mg²⁺ in place of Ba²⁺.</i>	1 1	
	ACE Interpretation	Gives appropriate evidence (minimum – as above) for identification of two of the ions. Evidence for Ca ²⁺ (from reaction with OH ⁻) may be allowed here	1	[3]
If a candidate does not record any practical work for sections (iii) and (iv), ions such as Ca ²⁺ , Mg ²⁺ or Zn ²⁺ may be credited from appropriate observations with NaOH and NH ₃ (aq)				
2 (e)	MMO Decisions	Selects CrO ₄ ²⁻ or Cr ₂ O ₇ ²⁻ as additional reagent (No cation or “aqueous ion” required)	1	[1]
			Qn 2 Total 15	