

CAMBRIDGE INTERNATIONAL EXAMINATIONS
GCE Advanced Subsidiary Level and GCE Advanced Level

MARK SCHEME for the May/June 2013 series

9701 CHEMISTRY

9701/23

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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1 (a) (i)



S atom has 6 **and** C atom has 4 electrons (1)

S=C double bonds (4 electrons) clearly shown (1)

(ii) linear **and** 180° (1) [3]

(b) (i) $\text{CS}_2 + 3\text{O}_2 \rightarrow \text{CO}_2 + 2\text{SO}_2$ (1)

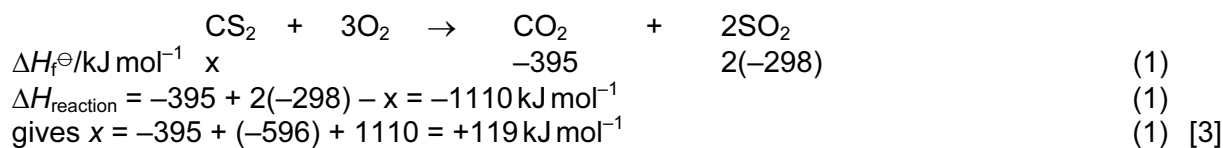
(ii) enthalpy change when 1 mol of a substance (1)

is burnt in an excess of oxygen/air

or is completely combusted

under standard conditions (1) [3]

(c)



(d) (i) $\text{CS}_2 + 2\text{NO} \rightarrow \text{CO}_2 + 2\text{S} + \text{N}_2$
or
 $\text{CS}_2 + 2\text{NO} \rightarrow \text{CO} + 2\text{S} + \text{N}_2\text{O}$

correct products (1)

correct equation (1)

(ii) from -2 to 0 **both** required (1) [3]

[Total: 12]

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- 2 (a) (i) if the conditions of a system in equilibrium are changed (1)
the position of equilibrium moves so as to reduce that change (1) [2]
- (ii) lower temperature (1)
because the forward reaction is exothermic (1)
higher pressure (1)
because the forward reaction shows a reduction in volume
or
there are fewer molecules/moles on RHS of equilibrium (1) [4]

(b)

	CO ₂	+	H ₂	⇌	CO	+	H ₂ O	
initial moles	0.70		0.70		0.30		0.30	
equil. moles	(0.70-x)		(0.70-x)		(0.30+x)		(0.30+x)	(1)
equil. concn.	$\frac{(0.70-x)}{1}$		$\frac{(0.70-x)}{1}$		$\frac{(0.30+x)}{1}$		$\frac{(0.30+x)}{1}$	

$$K_c = \frac{(0.30+x)^2}{(0.70-x)^2} = 1.44 \quad (1)$$

gives $x = 0.25$ (1)

at equilibrium,

$$n(\text{CO}_2) = n(\text{H}_2) = 0.70 - 0.25 = 0.45 \text{ moles}$$

and

$$n(\text{CO}) = n(\text{H}_2\text{O}) = 0.3 + 0.25 = 0.55 \text{ moles} \quad (1) \quad [4]$$

[Total: 10]

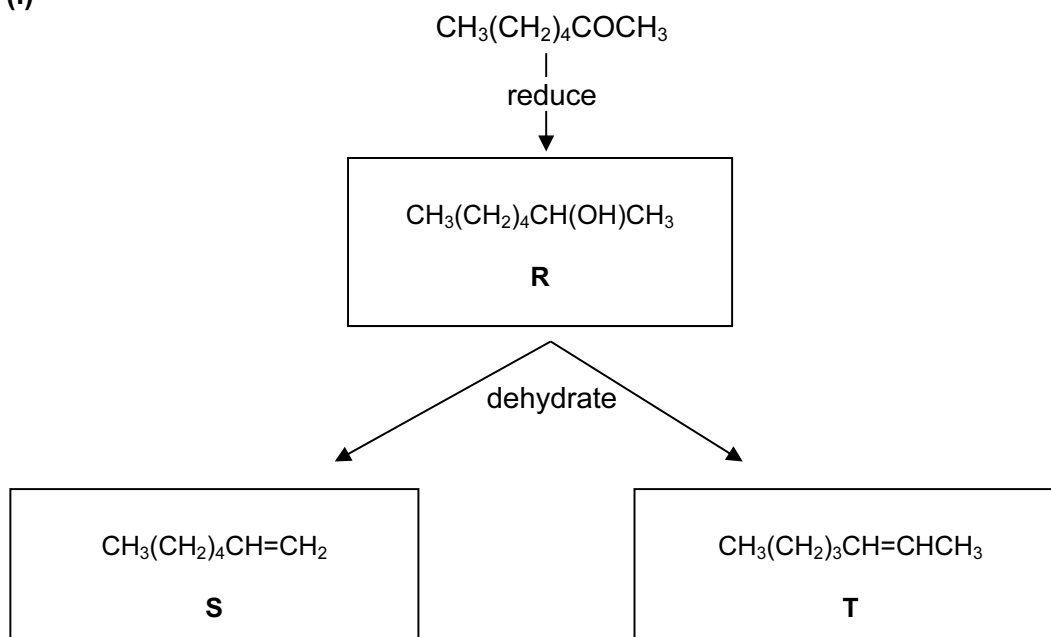
Page 4	Mark Scheme	Syllabus	Paper 1
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- 3 (a) (i) He or Ne or Ar or Kr (1)
- (ii) P or As (1)
- (iii) Br (1)
- (iv) Na allow Ar (1)
- (v) Si (1)
- (vi) P allow Si (1)
- (vii) Cl or F or Br (1) [7]
- (b) (i) any **two** from P_4O_6 , SO_2 and Cl_2O_7 (1+1)
- (ii) Al_2O_3 or SiO_2 (1)
- (iii) $MgSO_3$ (1) [4]
- (c) (i) Si is giant molecular/giant covalent **or**
P, S, and Cl are simple molecular (1)
- (ii) the molecules are S_8 , P_4 , Cl_2 (1)
- larger molecules have more electrons (1)
- and hence greater van der Waals' forces (1) [4]

[Total: 15]

Page 5	Mark Scheme	Syllabus	Paper 1
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4 (a) (i)



one mark for each correct compound, **R**, **S** and **T**

allow correct *cis* and *trans* versions of compound **T** for 2 marks (3 × 1)

(ii) reduction

NaBH_4 or LiAlH_4 or H_2/Ni or $\text{Na}/\text{C}_2\text{H}_5\text{OH}$ (1)

dehydration

$\text{P}_4\text{O}_{10}/\text{P}_2\text{O}_5$ or H_3PO_4 or conc. H_2SO_4 or Al_2O_3 (1) [5]

(b)

Tollens' reagent	NO REACTION
HCN	$ \begin{array}{c} \text{CH}_3(\text{CH}_2)_4\text{C}(\text{OH})\text{CH}_3 \\ \\ \text{CN} \end{array} $
$\text{K}_2\text{Cr}_2\text{O}_7/\text{H}^+$	NO REACTION

one mark for each correct answer (3 × 1) [3]

Page 6	Mark Scheme	Syllabus	Paper 1
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(c) Na_2CO_3 or NaHCO_3 effervescence/colourless gas

or

Na colourless gas

or

$\text{PCl}_3/\text{PCl}_5$ etc. steamy fumes

or

$\text{C}_2\text{H}_5\text{OH}/\text{conc. H}_2\text{SO}_4$ sweet smell of ester

or

$\text{K}_2\text{Cr}_2\text{O}_7/\text{H}^+$ orange solution becomes green

correct reagent

(1)

correct observation

(1) [2]

[Total: 10]

Page 7	Mark Scheme	Syllabus	Paper 1
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- 5 (a) (i) $\text{CH}_2=\text{CHCO}_2\text{H}$ (1)
- (ii) $\text{BrCH}_2\text{CHBrCH}_2\text{OH}$ (1)
- (iii) product is $\text{HOCH}_2\text{CH}(\text{OH})\text{CH}_2\text{OH}$
 correct addition across $>\text{C}=\text{C}<$ (1)
 original $-\text{CH}_2\text{OH}$ remains (1)
- (iv) $\text{HO}_2\text{CCO}_2\text{H}$ (1) [5]
- (b) (i) nucleophilic substitution (1)
- (ii) oxidation (1) [2]
- (c) (i) **step I**
 H_2 (1)
 heat with Ni catalyst (1)
- step II**
 acidified $\text{K}_2\text{Cr}_2\text{O}_7$ (1)
 heat **or** distil off product (1)
- (ii) structural isomerism
or
 functional group isomerism (1) [5]
- (d) **both** oxidation **and** reduction have occurred **or**
 disproportionation has taken place (1) [1]

[Total: 13]