

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Level and GCE Advanced Subsidiary Level

MARK SCHEME for the May/June 2006 question paper**9701 CHEMISTRY**

9701/04

Paper 4

Maximum raw mark 60

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which Examiners were initially instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

The minimum marks in these components needed for various grades were previously published with these mark schemes, but are now instead included in the Report on the Examination for this session.

- CIE will not enter into discussion or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the May/June 2006 question papers for most IGCSE and GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



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- 1 (a) (i) order w.r.t. NO = 2 [1]
order w.r.t. O₂ = 1 [1]
- (ii) rate = k[NO]²[O₂] [1]
- taking the first row: k = rate/([NO]²[O₂])
= 0.020/(0.005² × 0.005)
= **1.6 × 10⁵** ecf [1]
units = **mol⁻²dm⁶sec⁻¹** ecf [1]
- (iii) rate = k[NO]²[O₂]
= 1.6 × 10⁵ × 0.0025² × 0.0025
= **2.5 × 10⁻³** (mol dm⁻³ s⁻¹) ecf [1]
- [6]
- (b) (i) homogeneous [1]
- (ii) NO + ½ O₂ → NO₂
SO₂ + NO₂ → SO₃ + NO
(SO₃ + H₂O → H₂SO₄) [2]
- [3]
- [Total: 9]
- 2 (a) (i) K_{sp} = [Ca²⁺][SO₄²⁻] [1]
units are: mol²dm⁻⁶ ecf [1]
- (ii) [CaSO₄] = √K_{sp} = **5.5 × 10⁻³** (5.477 × 10⁻³)(mol dm⁻³) ecf [1]
- (iii) n(CaSO₄) in 100 dm³ = 5.5 × 10⁻³ × 100 = 0.55 moles ecf from (ii) [1]
M_r(CaSO₄) = 136.1
Thus mass(CaSO₄) = 0.55 × 136.1 = **74.8g** (0.55 × M_r) [1]
(if the accurate [CaSO₄] is held throughout the calculation, ans = **74.5g**) [5]
- (b) (i) down the group: the ΔH_{solution} becomes more endothermic;
both lattice energy **and** ΔH_{hydration} become less (exothermic);
due to **ionic** radius (of M²⁺) increasing;
but ΔH_{hydration} changes more than lattice energy *any three points* [3]
- (ii) K_{sp} = [Ba²⁺][SO₄²⁻] = (9 × 10⁻⁶)² = **8.1 × 10⁻¹¹** **NO** ecf [1]
- [4]
- (c) (i) LE is the energy change when **1 mole** of (ionic) solid [1]
is formed from its **gaseous ions** [1]
- (ii) LE(BaSO₄) < LE(MgSO₄), due to larger radius of Ba²⁺ both points [1]
- [3]
- [Total: 12]

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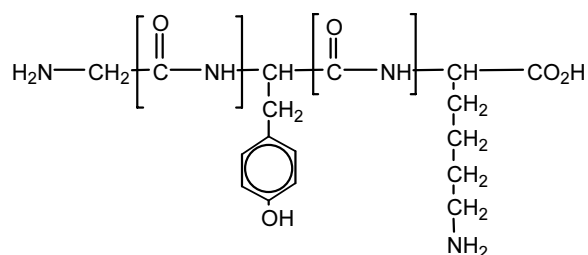
- 3 (a) (i) +2, +3, +4, +5 (ignore 0 and +1) all four [1]
- (ii) [Ar]3d² [1]
- [2]**
- (b) (i) take a fixed amount/aliquot/pipette-full of the Fe²⁺ solution [1]
- titrate with KMnO₄ in **the burette** [1]
- until the first permanent pink colour (*or* change from colourless to pink) [1]
- repeat until two titres are within 0.1 cm³ [1]
- MnO₄⁻ + 8H⁺ + 5Fe²⁺ → Mn²⁺ + 4H₂O + 5Fe³⁺ [1]
 (*or* molecular equn.)
- (ii) n(MnO₄⁻) = 0.02 x 14/1000 = 2.8 x 10⁻⁴ moles [1]
- n(Fe²⁺) in 25 cm³ = 2.8 x 10⁻⁴ x 5 (x 5) [1]
 (= 1.4 x 10⁻³ moles)
- n(Fe²⁺) in 100 cm³ = 1.4 x 10⁻³ x 4 (x 4) [1]
 (= 5.6 x 10⁻³ moles)
- mass of Fe in 2.0 g ore = 5.6 x 10⁻³ x 55.8
 = 0.31 g
- percentage = 100 x 0.31/2 = **15.6%** (*use of 55.8 or 56 and %*) [1]
- [9]**
- (c) (i) Cu²⁺(aq) or [Cu(H₂O)₆]²⁺ [1]
- (ii) **pale blue ppt.** (of Cu(OH)₂(s)) [*ignore any refs. to iron hydroxides*] [1]
 (which dissolves to give....)
- a deep blue solution** [1]
- which contains [Cu(NH₃)₄]²⁺ ions (*can be read into equn, below*) [1]
- formed by ligand displacement [1]
or an equation such as
- Cu(OH)₂ + 4NH₃ → [Cu(H₂O)₄]²⁺ + 2OH⁻
or [Cu(H₂O)₆]²⁺ + 4NH₃ → [Cu(H₂O)₄]²⁺ + 6H₂O
- [5]**

[Total: 16 max 14]

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- 4 (a) HCl or H_2SO_4 or H^+ or acid [1]
 conc(if HCl only)/dilute/aqueous + heat [1] [2]
- (b) **two** rings only (1 ring around the α -C of tyrosine & 1 around the α -C of lysine) [1] [1]
- (c) $^+\text{NH}_3\text{CH}_2\text{CO}_2^-$ (or displayed formula) [1] [1]
- (d) (i) $\text{NH}_2\text{CH}_2\text{CO}_2^- (\text{Na}^+)$ (either $-\text{CO}_2^-\text{Na}^+$ or $-\text{CO}_2\text{Na}$ but NOT $-\text{CO}-\text{O}-\text{Na}$) [1]
 (ii) $(\text{Na}^+) ^-\text{O}-\text{C}_6\text{H}_4-\text{CH}_2\text{CH}(\text{NH}_2)\text{CO}_2^- (\text{Na}^+)$ [1] + [1]
 (iii) $(\text{Cl})^+\text{NH}_3(\text{CH}_2)_4\text{CH}(\text{NH}_3^+)\text{CO}_2\text{H} (\text{Cl})$ [1] + [1]
 (iv) $\text{HO}-\text{C}_6\text{H}_2\text{Br}_2-\text{CH}_2\text{CH}(\text{NH}_2)\text{CO}_2\text{H}$ (if shown, Br at 2,6 to OH group) [1] [6]

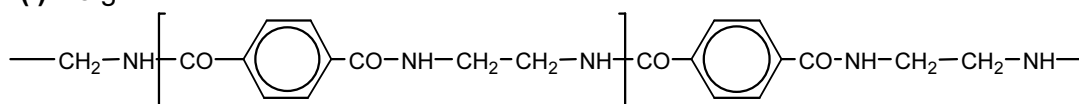
(e)



structure [1]
 at least one peptide group identified [1]

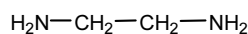
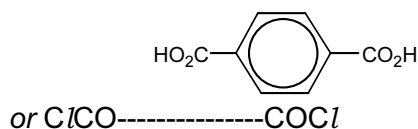
[2]

(f) (i) e.g.



[1]

(ii)



[1]

[1]

[3]

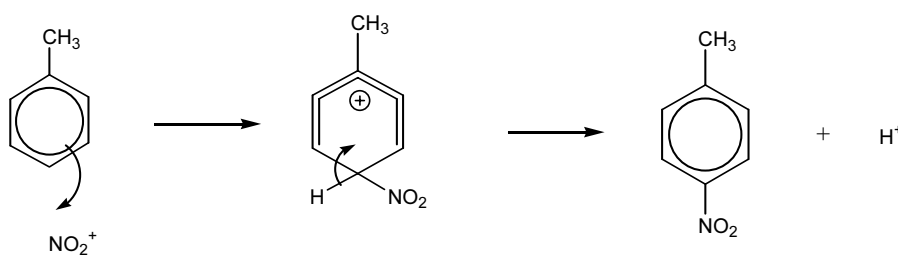
[Total: 15 max 14]

Page 4	Mark Scheme	Syllabus	Paper 1
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- 5 (a) I: $\text{HNO}_3 + \text{H}_2\text{SO}_4$ (or names) [1]
 (both) conc. and at $50^\circ\text{C} < T < 60^\circ\text{C}$ ✓ [1]
 II: $\text{KMnO}_4 (+\text{OH}^-) + \text{heat}$ [1]
 III: $\text{Sn} + (\text{conc}) \text{HCl}$ [1]
 IV: $\text{CH}_3\text{CH}_2\text{OH}$ (or name) [1]
 + c. $\text{H}_2\text{SO}_4 + \text{heat}$ [1]

[6]

(b)



- intermediate, including \oplus [1]
 NO_2^+ at start and H^+ at finish [1]
 (no marks for curly arrows, but if present, they must be in correct direction)

[2]

- (c) (i) ester and (primary) amine [2]
 (ii) **more basic**: amine group is *not* adjacent to benzene ring **both** points [1]
 (or lone pair (on N) is not delocalised)

[3]

[Total: 11]